

**GCSE Double science (Chemistry booklet) - 11W2**  
**& 11E3**

As you have hopefully seen in your lesson, the Double award GCSE is made up on the following 8 components each worth 25% of your GCSE:

C1 Exam - 1 Hour (studied in Year 10)

B1 Exam - 1 Hour (studied in Year 10)

P1 Exam - 1 hour (Studied in Year 10)

CA - COMPLETED

= Core Science GCSE

C2 Exam - 1 Hour (Currently studying)

B2 Exam - 1 Hour (Currently studying)

P2 Exam - 1 hour (Currently studying)

CA - COMPLETED

= Additional Science GCSE

You mock exams are approaching fast and you will be completing the following examinations:

- C1 examination (1 hour)
- C2 examination (1 hour)

If you are looking for support with revision, there are the following options available to all:

1. Monday Lunch time club - Mr Bryan Science 3
2. After-school Wednesday revision - various rooms
3. GCSE revision guides - See Mr Bryan if you still need one
4. BBC bitesize
5. Lots of revision materials are on KEYS.
6. See your teacher at any time!

The following pages have the criteria that can be assessed in the examinations, I recommend using the tick boxes in your revision, so that you are thorough and cover everything.

# C1 learning outcomes:

## Topic 1 The Earth's sea and atmosphere

- 1.1 Recall that the gases produced by volcanic activity formed the Earth's early atmosphere
- 1.2 Recall that the early atmosphere contained: a little or no oxygen b a large amount of carbon dioxide c water vapour and small amounts of other gases
- 1.3 Explain why there are different sources of information about the development of the atmosphere which makes it difficult to be precise about the evolution of the atmosphere
- 1.4 Describe how condensation of water vapour formed oceans
- 1.5 Describe how the amount of carbon dioxide in the atmosphere was reduced by: a the dissolution of carbon dioxide into the oceans b the later incorporation of this dissolved carbon dioxide into marine organisms which eventually formed carbonate rocks
- 1.6 Explain how the growth of primitive plants used carbon dioxide and released oxygen by photosynthesis and consequently the amount of oxygen in the atmosphere gradually increased
- 1.7 Investigate the proportion of oxygen in the atmosphere
- 1.8 Describe the current composition of the atmosphere and interpret data sources showing this information
- 1.9 Demonstrate an understanding of how small changes in the atmosphere occur through: a volcanic activity b human activity, including the burning of fossil fuels, farming and deforestation

## Topic 2

- 2.1 Describe that igneous rocks, such as granite, are: a formed by the solidification of magma or lava b made of crystals whose size depends on the rate of cooling
- 2.2 Describe chalk and limestone as examples of sedimentary rocks
- 2.3 Describe how sedimentary rocks are formed by the compaction of layers of sediment over a very long time period
- 2.4 Recall that sedimentary rocks: a may contain fossils b are susceptible to erosion

2.5 Describe marble as an example of a metamorphic rock

2.6 Describe the formation of metamorphic rocks by the action of heat and/or pressure, including the formation of marble from chalk or limestone

2.7 Recall that limestone, chalk and marble exist in the Earth's crust and that they are all natural forms of calcium carbonate

2.8 Demonstrate an understanding of the balance between the demand for limestone and the economic, environmental and social effects of quarrying it

2.9 Demonstrate an understanding of the commercial need for quarrying calcium carbonate on a large scale, as a raw material, for the formation of glass, cement and concrete

2.10 Describe the thermal decomposition of calcium carbonate into calcium oxide and carbon dioxide

2.11 Investigate the ease of thermal decomposition of carbonates, including calcium carbonate, zinc carbonate and copper carbonate

2.12 Describe the ease of thermal decomposition of different metal carbonates

2.13 Demonstrate an understanding that: a atoms are the smallest particles of an element that can take part in chemical reactions b during chemical reactions, atoms are neither created nor destroyed c during chemical reactions, atoms are rearranged to make new products with different properties from the reactants

2.14 Describe the effect of water on calcium oxide

2.15 Describe how calcium hydroxide dissolves in water to form a solution, known as limewater

2.16 Demonstrate an understanding that the total mass before and after a reaction in a sealed container is unchanged, as shown practically by a precipitation reaction

2.17 Explain how calcium oxide, calcium hydroxide and calcium carbonate can be used to neutralise soil acidity

2.18 Explain how calcium carbonate can be used to remove acidic gases from coal-fired power station chimneys, reducing harmful emissions and helping to reduce acid rain

### Topic 3 Acids

3.1 Recall that hydrochloric acid is produced in the stomach to: a help digestion b kill bacteria

3.2 Describe indigestion remedies as containing substances that neutralise excess stomach acid

3.3 Investigate the effectiveness of different indigestion remedies

3.4 Recall that acids are neutralised by: a metal oxides b metal hydroxides c metal carbonates to produce salts (no details of salt preparation techniques or ions are required)

3.5 Recall that: a hydrochloric acid produces chloride salts b nitric acid produces nitrate salts c sulfuric acid produces sulfate salts

3.6 Describe electrolysis as a process in which electrical energy, from a d.c. supply, decomposes compounds, by considering the electrolysis of dilute hydrochloric acid to produce hydrogen and chlorine (explanations of the reactions at the electrodes are not required)

3.7 Investigate the electrolysis of dilute hydrochloric acid

3.8 Describe the chemical test for hydrogen

3.9 Describe the chemical test for chlorine Recall that chlorine can be obtained from sea water by electrolysis (explanations of the reactions at the electrodes are not required)

3.11 Describe chlorine as a toxic gas and that this leads to potential hazards associated with its large-scale manufacture

3.12 Describe the use of chlorine in the manufacture of bleach and of the polymer poly(chloroethene) (PVC)

3.13 Recall that water can be decomposed by electrolysis to form hydrogen and oxygen

3.14 Describe the chemical test for oxygen

#### Topic 4 Obtaining and using metals

4.1 Recall that: a most metals are extracted from ores found in the Earth's crust b unreactive metals are found in the Earth as the uncombined elements

4.2 Describe how most metals are extracted from their ores by: a heating with carbon, illustrated by iron b electrolysis, illustrated by aluminium (knowledge of the blast furnace or the electrolytic cell for aluminium extraction are not required)

4.3 Explain why the method used to extract a metal is related to its position in the reactivity series and cost of the extraction process

- 4.4 Investigate methods for extracting a metal from its ore
- 4.5 Describe oxidation as the gain of oxygen and reduction as the loss of oxygen
- 4.6 Recall that the extraction of metals involves reduction of ores
- 4.7 Recall that the oxidation of metals results in corrosion
- 4.8 Demonstrate an understanding that a metal's resistance to oxidation is related to its position in the reactivity series
- 4.9 Discuss the advantages of recycling metals, including economic implications, and how recycling preserves both the environment and the supply of valuable raw materials
- 4.10 Describe the uses of metals in relation to their properties, including: a aluminium b copper c gold d steel
- 4.11 Use models to explain why converting pure metals into alloys often increases the strength of the product
- 4.12 Demonstrate an understanding that iron is alloyed with other metals to produce alloy steels with a higher strength and a greater resistance to corrosion
- 4.13 Describe how alloying changes the properties of metals, including: a smart or shape memory alloys, including nitinol, an alloy of nickel and titanium b gold alloys with higher strength, including fineness (parts per thousand) and carats to indicate the proportion of pure gold
- 4.14 Demonstrate an understanding that new materials are developed by chemists to fit new applications, such as the creation of new shape memory alloys for use, for example, in spectacle frames and as stents in damaged blood vessels

## Topic 5 Fuels

- 5.1 Describe hydrocarbons as compounds that contain carbon and hydrogen only
- 5.2 Describe crude oil as a complex mixture of hydrocarbons
- 5.3 Describe the separation of crude oil into simpler, more useful mixtures by the process of fractional distillation (details of fractional distillation are not required)
- 5.4 Recall the name and uses of the following fractions: a gases, used in domestic heating and cooking b petrol, used as fuel for cars c kerosene, used as fuel for aircraft d diesel oil, used as fuel for some cars and trains e fuel oil, used as fuel for large ships and in some power stations f bitumen, used to surface roads and roofs

5.5 Describe that hydrocarbons in different fractions differ from each other in: a the number of carbon and hydrogen atoms their molecules contain b boiling points c ease of ignition d viscosity

5.6 Describe how the complete combustion of hydrocarbons: a involves the oxidation of the hydrocarbons b produces carbon dioxide and water c gives out energy

5.7 Describe the chemical test for carbon dioxide (using limewater)

5.8 Explain why the incomplete combustion of hydrocarbons can produce carbon and carbon monoxide

5.9 Describe how carbon monoxide behaves as a toxic gas

5.10 Demonstrate an understanding of the problems caused by incomplete combustion producing carbon monoxide and soot in appliances that use carbon compounds as fuels

5.11 Explain why impurities in some hydrocarbon fuels result in the production of sulfur dioxide

5.12 Demonstrate an understanding of some problems associated with acid rain caused when sulfur dioxide dissolves in rain water

5.13 Describe how various gases in the atmosphere, including carbon dioxide, methane and water vapour, trap heat from the Sun and that this keeps the Earth warm

5.14 Demonstrate an understanding that the Earth's temperature varies and that human activity may influence this

5.15 Demonstrate an understanding that the proportion of carbon dioxide in the atmosphere varies, due to human activity, and that chemists are investigating methods to control the amount of the gas in the atmosphere by: a iron seeding of oceans b converting carbon dioxide into hydrocarbons

5.16 Evaluate how far the correlation between global temperature and the proportion of carbon dioxide in the atmosphere provides evidence for climate change

5.17 Describe biofuels as being possible alternatives to fossil fuels

5.18 Recall that one example of a biofuel is ethanol obtained by processing sugar cane or sugar beet and that it can be used to reduce the demand for petrol

5.19 Evaluate the advantages and disadvantages of replacing fossil fuels with biofuels, including: a the fact that biofuels are renewable b that growing the crops to make biofuels requires land and may affect the availability of land for growing food c the

balance between the carbon dioxide removed from the atmosphere as these crops grow and the carbon dioxide produced when they are transported and burned

5.20 Demonstrate an understanding of the factors that make a good fuel, including: a how easily it burns b the amount of ash or smoke it produces c the comparative amount of heat energy it produces (calculations involving conversion to joules are not required) d how easy it is to store and transport

5.21 Recall that a simple fuel cell combines hydrogen and oxygen to form water and that this reaction releases energy

5.22 Evaluate the advantages and disadvantages of using hydrogen, rather than petrol, as a fuel in cars

5.23 Describe petrol, kerosene and diesel oil as non-renewable fossil fuels obtained from crude oil and methane as a non-renewable fossil fuel found in natural gas

5.24 Compare the temperature rise produced when the same volume of water is heated by different fuels

5.25 Recall that alkanes are saturated hydrocarbons, which are present in crude oil

5.26 Recall the formulae of the alkanes methane, ethane and propane, and draw the structures of these molecules to show how the atoms are bonded together (no further knowledge of bonding is required in this unit)

5.27 Recall that alkenes are unsaturated hydrocarbons

5.28 Recall the formulae of the alkenes ethene and propene and draw the structures of their molecules to show how the atoms are bonded together (No further knowledge of bonding is required in this unit)

5.29 Describe how bromine water is used to distinguish between alkanes and alkenes

5.30 Describe how cracking involves the breaking down of larger saturated hydrocarbon molecules (alkanes) into smaller, more useful ones, some of which are unsaturated (alkenes)

5.31 Explain why cracking is necessary, including by using data on the composition of different crude oils and the demand for fractions in crude oil

5.32 Describe the cracking of liquid paraffin in the laboratory



5.33 Recall that: a many ethene molecules can combine together in a polymerisation reaction b the polymer formed is called poly(ethene) (conditions and mechanisms not required but equations required)

5.34 Describe how other polymers can be made by combining together other monomer molecules, to include poly(propene), poly(chloroethene) (PVC) and PTFE

5.35 Relate uses of the polymers poly(ethene), poly(propene), poly(chloroethene) (PVC) and PTFE to the properties of the compounds

5.36 Recall that most polymers are not biodegradable, that they persist in landfill sites and that many produce toxic products when burnt

5.37 Explain how some problems associated with the disposal of polymers can be overcome: a by recycling b by developing biodegradable polymers

## C2 Learning outcomes:

<u>C2 Atomic Structure &amp; the Periodic Table</u> <i>Italic text</i> = practical task <b>Bold text</b> = higher tier only	Start RAG	End RAG	Revised 1 ✓	Revised 2 ✓
0.1 Recall the formulae of elements and simple compounds in the unit				
0.2 Represent chemical reactions by word equations and simple balanced equations				
0.3 Write balanced chemical equations including the use of state symbols (s), (l), (g) and (aq) for a wide range of reactions in this unit				
0.4 Assess practical work for risks and suggest suitable precautions for a range of practical scenarios for reactions in this unit				
0.5 Demonstrate an understanding that hazard symbols used on containers:				
a) indicate the dangers associated with the contents				
b) inform people about safe-working procedures with these substances in the laboratory				

<b>Topic 1: C2 Atomic Structure &amp; the Periodic Table</b>		Start RAG	End RAG	Revised 1 ✓	Revised 2 ✓
<i>Italic text</i> = practical task <b>Bold text</b> = higher tier only					
1.1	Explain how Mendeleev:				
	a. arranged the elements, known at that time, in a periodic table by using properties of these elements and their compounds				
	b. used his table to predict the existence and properties of some elements not then discovered				
1.2	Classify elements as metals or non-metals according to their position in the periodic table				
1.3	Describe the structure of an atom as a nucleus containing protons and neutrons, surrounded by electrons in shells (energy levels)				
1.4	Demonstrate an understanding that the nucleus of an atom is very small compared to the overall size of the atom				
1.5	Describe atoms of a given element as having the same number of protons in the nucleus and that this number is unique to that element				
1.6	Recall the relative charge and relative mass of:				
	a. proton				
	b. a neutron				
	c. an electron				
1.7	Demonstrate an understanding that atoms contain equal numbers of protons and electrons				
1.8	Explain the meaning of the terms:				
	a. atomic number				
	b. mass number				
	c. relative atomic mass				
1.9	Describe the arrangement of elements in the periodic table such that:				
	a. elements are arranged in order of increasing atomic number, in rows called periods				
	b. elements with similar properties are placed in the same vertical column, called groups				
1.10	<b>Demonstrate an understanding that the existence of isotopes results in some relative atomic masses not being whole numbers</b>				

I.11	<b>Calculate the relative atomic mass of an element from the relative masses and abundances of its isotopes</b>				
I.12	Apply rules about the filling of electron shells (energy levels) to predict the electronic configurations of the first 20 elements in the periodic table as diagrams and in the form 2.8.1				
I.13	Describe the connection between the number of outer electrons and the position of an element in the periodic table				

<p style="text-align: center;"><b>Topic 2: C2 Ionic Compounds &amp; Analysis</b></p> <p style="text-align: center;"><i>Italic text</i> = practical task <b>Bold text</b> = higher tier only</p>	Start RAG	End RAG	Revised 1 ✓	Revised 2 ✓
2.1 Demonstrate an understanding that atoms of different elements can combine to form compounds by the formation of new chemical bonds				
2.2 Describe how ionic bonds are formed by the transfer of electrons to produce cations and anions				
2.3 Describe an ion as an atom or group of atoms with a positive or negative charge				
2.4 Describe the formation of sodium ions, Na <sup>+</sup> , and chloride ions, Cl <sup>-</sup> , and hence the formation of ions in other ionic compounds from their atoms, limited to compounds of elements in groups 1, 2, 6 and 7				
2.5 Demonstrate an understanding of the use of the endings –ide and –ate in the names of compounds				
2.6 Deduce the formulae of ionic compounds (including oxides, hydroxides, halides, nitrates, carbonates and sulfates) given the formulae of the constituent ions				
<p>2.7 <b>Describe the structure of ionic compounds as a lattice structure:</b></p> <p style="padding-left: 20px;">a) <b>consisting of a regular arrangement of ions</b></p> <p style="padding-left: 20px;">b) <b>held together by strong electrostatic forces (ionic bonds) between oppositely-charged ions</b></p>				
2.8 Describe <b>and explain</b> the properties of ionic substances including sodium chloride and magnesium oxide, limited to:				
a) melting points and boiling points				
b) whether they conduct electricity as solids, when molten and in aqueous solution				
2.9 Recall the general rules which describe the solubility of common types of substances in water:				
a) all common sodium, potassium and ammonium salts are soluble				
b) all nitrates are soluble				
c) common chlorides are soluble except those of silver and lead				
d) common sulfates are soluble except those of lead, barium and calcium				
e) common carbonates and hydroxides are insoluble except those of sodium, potassium and ammonium				
2.10 Demonstrate an understanding that insoluble salts can be				

formed as precipitates by the reaction of suitable reagents in solution				
2.11 Demonstrate an understanding of the method needed to prepare a pure, dry sample of an insoluble salt				
2.12 <i>Prepare an insoluble salt by precipitation</i>				
2.13 Use solubility rules to predict whether a precipitate is formed when named solutions are mixed together and to name the precipitate				
2.14 Recall that the insoluble salt, barium sulfate, is given as a 'barium meal' to X-ray patients because				
a) it is opaque to X-rays				
b) it is safe to use as, although barium salts are toxic, its insolubility prevents it entering the blood				
2.15 Describe tests to show the following ions are present in solids or solutions:				
a) Na <sup>+</sup> , K <sup>+</sup> , Ca <sup>2+</sup> , Cu <sup>2+</sup> using flame tests				
b) CO <sub>3</sub> <sup>2-</sup> using dilute acid and identifying the carbon dioxide evolved				
c) SO <sub>4</sub> <sup>2-</sup> using dilute hydrochloric acid and barium chloride solution				
d) Cl <sup>-</sup> using dilute nitric acid and silver nitrate solution				
2.16 Recall that chemists use spectroscopy (a type of flame test) to detect the presence of very small amounts of elements and that this led to the discovery of new elements, including rubidium and caesium				

<b>Topic 3: C2 Covalent Compounds and Separation Techniques</b> <i>Italic text = practical task</i> <b>Bold text = higher tier only</b>	Start RAG	End RAG	Revised 1 ✓	Revised 2 ✓
3.1 Describe a covalent bond as a pair of electrons shared between two atoms				
3.2 Recall that covalent bonding results in the formation of molecules				
3.3 Explain the formation of simple molecular, covalent substances using dot and cross diagrams, including:				
a) hydrogen				
b) hydrogen chloride				
c) water				
d) methane				
e) <b>oxygen</b>				
f) <b>carbon dioxide</b>				
3.4 <i>Classify different types of elements and compounds by investigating their melting points and boiling points, solubility in water and electrical conductivity (as solids and in solution) including sodium chloride, magnesium sulfate, hexane, liquid paraffin, silicon(IV) oxide, copper sulfate, and sucrose (sugar)</i>				
3.5 Describe the properties of typical simple molecular, covalent compounds, limited to:				
a) low melting points and boiling points, in terms of weak forces between molecules				
b) poor conduction of electricity				
3.6 Demonstrate an understanding of the differences between the properties of simple molecular, covalent substances and those of giant molecular, covalent substances, including diamond and graphite				
3.7 <b>Explain why, although they are both forms of carbon and giant molecular substances, graphite is used to make electrodes and as a lubricant, whereas diamond is used in cutting tools</b>				
3.8 Describe the separation of two immiscible liquids using a separating funnel				
3.9 Describe the separation of mixtures of miscible liquids by fractional distillation, by referring to the fractional distillation of liquid air to produce nitrogen and oxygen				
3.10 Describe how paper chromatography can be used to separate and identify components of mixtures, including colouring agents in foodstuffs				

3.11 Evaluate the information provided by paper chromatograms, including the calculation of Rf values, in a variety of contexts, such as the food industry and forensic science

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<p style="text-align: center;"><b>Topic 4: C2 Groups in the Periodic Table</b></p> <p style="text-align: center;"><i>Italic text</i> = practical task <b>Bold text</b> = higher tier only</p>	Start RAG	End RAG	Revised 1 ✓	Revised 2 ✓
4.1 Classify elements as alkali metals (group 1), halogens (group 7), noble gases (group 0) and transition metals based on their position in the periodic table				
4.2 Describe the structure of metals as a regular arrangement of positive ions surrounded by a sea of delocalised electrons				
4.3 Describe and explain the properties of metals, limited to malleability and the ability to conduct electricity				
4.4 Recall that most metals are transition metals and that their typical properties include:				
a high melting point				
b the formation of coloured compounds				
4.5 Demonstrate an understanding that elements and compounds can be classified as				
a) ionic				
b) simple molecular covalent				
c) giant molecular covalent				
d) metallic				
and that each type of substance has different physical properties, including relative melting point and boiling point, relative solubility in water and ability to conduct electricity (as solids and in solution)				
4.6 Describe alkali metals as				
a) soft metals				
b) metals with comparatively low melting points				
4.7 Describe the reactions of lithium, sodium and potassium with water to form hydroxides which are alkaline, and hydrogen gas				
4.8 Describe the pattern in reactivity of the alkali metals lithium, sodium and potassium with water and use this pattern to predict the reactivity of other alkali metals <b>and explain the pattern</b>				
4.9 Recall the colours and physical states of the halogens at room temperature				
4.10 Describe the reaction of halogens with metals to form metal halides				
4.11 Recall that halogens react with hydrogen to produce hydrogen halides which dissolve in water to form acidic solutions				
4.12 <i>Investigate displacement reactions of halogens reacting with halide</i>				



<i>ions in solution</i>				
4.13 Describe the relative reactivity of the halogens as shown by their displacement reactions with halide ions in aqueous solution				
4.14 Describe the noble gases as chemically inert, compared with the other elements and demonstrate an understanding that this lack of reactivity can be explained by the electronic arrangements in their atoms				
4.15 Demonstrate an understanding that the discovery of the noble gases was due to chemists:				
a) noticing that the density of nitrogen made in a reaction differed from that of nitrogen obtained from air				
b) developing a hypothesis about the composition of the air				
c) performing experiments to test this hypothesis and show the presence of the noble gases				
4.16 Relate the uses of the noble gases to their properties, including:				
a) inertness (including providing an inert atmosphere for welding and in filament lamps)				
b) low density (including filling balloons)				
c) non-flammability				
4.17 Use the pattern in a physical property of the noble gases, such as boiling point or density, to estimate an unknown value for another member of the group				

<p style="text-align: center;"><b>Topic 5: C2 Chemical Reactions</b></p> <p style="text-align: center;"><i>Italic text = practical task</i> <b>Bold text = higher tier only</b></p>	Start RAG	End RAG	Revised 1 ✓	Revised 2 ✓
5.1 <i>Measure temperature changes accompanying some of the following types of change:</i>				
a) <i>salts dissolving in water</i>				
b) <i>neutralisation reactions</i>				
c) <i>displacement reactions</i>				
d) <i>precipitation reactions</i>				
5.2 Define an exothermic change or reaction as one in which heat energy is given out, including combustion reactions or explosions				
5.3 Define an endothermic change or reaction as one in which heat energy is taken in, including photosynthesis or dissolving ammonium nitrate in water				
5.4 Describe the breaking of bonds as endothermic and the making of bonds as exothermic				
5.5 Demonstrate an understanding that the overall heat energy change for a reaction is				
a) exothermic if more heat energy is released making bonds in the products than is required to break bonds in the reactants				
b) endothermic if less heat energy is released making bonds in the products than is required to break bonds in the reactants				
<b>5.6 Draw and interpret simple graphical representations of energy changes occurring in chemical reactions (no knowledge of activation energy is required)</b>				
5.7 <i>Investigate the effect of temperature, concentration and surface area of a solid on the rate of a reaction such as hydrochloric acid and marble chips</i>				
5.8 Recall that the rates of chemical reactions vary from very fast, explosive reactions to very slow reactions				
5.9 Describe the effect of changes in temperature, concentration <b>and</b> surface area of a solid on the rate of reaction				
5.10 Describe how reactions can occur when particles collide <b>and explain how rates of reaction are increased by increasing the frequency and/or energy of collisions</b>				
<b>5.11 Demonstrate an understanding that not all collisions lead to a reaction, especially if particles collide with low energy</b>				

5.12 Recall the effect of a catalyst on the rate of reaction				
5.13 Demonstrate an understanding that catalytic converters in cars:				
a) have a high surface area, to increase the rate of reaction of carbon monoxide and unburnt fuel from exhaust gases with oxygen from the air to produce carbon dioxide and water				
b) work best at high temperatures				

<p style="text-align: center;"><b>Topic 6: C2 Quantitative Chemistry</b></p> <p style="text-align: center;"><i>Italic text</i> = practical task <b>Bold text</b> = higher tier only</p>	Start RAG	End RAG	Revised 1 ✓	Revised 2 ✓
6.1 Calculate relative formula mass given relative atomic masses				
6.2 Calculate the formulae of simple compounds from reacting masses and understand that these are empirical formulae				
6.3 <i>Determine the empirical formula of a simple compound, such as magnesium oxide</i>				
6.4 Calculate the percentage composition by mass of a compound from its formula and the relative atomic masses of its constituent elements				
<b>6.5 Use balanced equations to calculate masses of reactants and products</b>				
6.6 Recall that the yield of a reaction is the mass of product obtained in the reaction				
6.7 Demonstrate an understanding that the actual yield of a reaction is usually less than the yield calculated using the chemical equation (theoretical yield)				
6.8 Calculate the percentage yield of a reaction from the actual yield and the theoretical yield				
6.9 Demonstrate an understanding of the reasons why reactions do not give the theoretical yield due to factors, including:				
a) incomplete reactions				
b) practical losses during the preparation				
c) competing, unwanted reactions				
6.10 Demonstrate an understanding that many reactions produce waste products which:				
a) are not commercially useful				
b) can present economic, environmental and social problems for disposal				
<b>6.11 Demonstrate an understanding that chemists in industry work to find the economically most favourable reactions where</b>				
<b>a) the percentage yield is high</b>				
<b>b) the products of the reaction are commercially useful</b>				
<b>c) the reaction occurs at a suitable speed</b>				

